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CBSE 10th Acid Bases and Salts Unsolved Paper

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CBSE 10th Acid Bases and Salts

Unsolved Paper

SECTION - A

Q.1. which of the following gives CO_2 on heating?

- (a) Slaked
- (b) Quick lime
- (c) Lime stone
- (d) Soda ash.

Q.2. which is a base and not alkali?

- (a) NaOH
- (b) KOH
- (c) $\text{Fe}(\text{OH})_3$
- (d) None

Q.3. Plaster of Paris is made from

- (a) Lime stone
- (b) Slaked Lime
- (c) Quick lime
- (d) Gypsum

Q.4. An aqueous solution with pH-zero is

- (a) Acidic
- (b) Alkaline
- (c) Neutral
- (d) Amphoteric

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Q.5. The H^+ ion concentration of a solution is $1.0 \times 10^{-5}m$. The solution is

- (a) Acidic
- (b) Alkaline
- (c) Neutral
- (d) Amphoteric

Q.6. The difference of water molecules is gypsum and Plaster of Paris is

- (a) $\frac{5}{2}$
- (b) 2
- (c) $\frac{1}{2}$
- (d) $\frac{3}{2}$

Q.7. Setting of Plaster of Paris takes place due to

- (a) Oxidation
- (b) Reduction
- (c) Dehydration
- (d) Hydration

Q.8. The odour of acetic acid resembles that of

- (a) Rose
- (b) Burning Plastic
- (c) Vinegar
- (d) Kerosene

Q.9. Plaster of Paris hardens by

- (a) Giving off CO_2
- (b) Changing into $CaCO_3$
- (c) Combining with water
- (d) Giving out water

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Q.10. A drop of liquid sample was put on the pH paper, paper turned blue. The liquid sample must be of

- (a) Lemon Juice
- (b) HCl
- (c) Sodium bicarbonate
- (d) Ethanoic acid.

SECTION – B

Q.11. What happens to the crystals of washing soda when exposed to air?

Q.12. Why is sodium hydrogen carbonate an essential ingredient in antacids?

Q.13. What is the chemical name of washing soda? Name three raw materials used in making washing soda by Solvay process?

Q.14. Why should curd and sour substance not be kept in brass and copper vessels.

Q.15. While diluting an acid, why it is recommended that the acid should be added to water and not water to the acid?

Q.16. Why does distilled water not conduct electricity, whereas rain water does?

Q.17. Why does dry HCl gas not change the colour of the dry litmus paper?

SECTION - C

Q.18. (a) Name the raw materials used in the manufacture of sodium carbonate by Solvay process?

(b) How is sodium hydrogen carbonate formed from a mixture of NH_4Cl and $NaHCO_3$?

Q.19. What will you observe when:

(i) Red litmus is introduced into a solution of sodium sulphate.

(ii) Methyl orange is added to dil HCl.

(iii). Blue litmus is introduced into a solution of ferric chloride

Q.20. Explain why?

(a) Common salt becomes sticky during the rainy season.

(b) Blue vitriol changes to white upon heating.

Q.21. Metal compound 'A' reacts with dilute hydrochloric acid to produce effervescence. The gas evolved extinguishes a burning candle. Write a balanced chemical equation for the reaction, if one of the compounds formed is calcium chloride.

Q.22. A compound X of sodium is commonly used in kitchen for making crispy pakoras. It is also used for curing acidity in the stomach. Identify 'X'. What is its chemical formula? State the reaction that takes place when it is heated during cooking?

Q.23. Explain why-

(i) Anhydrous calcium chloride is used in desiccators

(ii) If a bottle full of concentrated H_2SO_4 is left open in the atmosphere by accident, the acid starts flowing out of the bottle of its own.

Q.24. A first aid manual suggests that vinegar should be used to treat wasp sting and baking soda for bee stings.

- (a) What does this information tell you about the chemical name of the wasp sting?
- (b) If there were no baking soda in the house, what other house hold substances would you use to treat as stings?

SECTION – D

Q.25. (a) Why does an aqueous solution of acid conduct electricity?

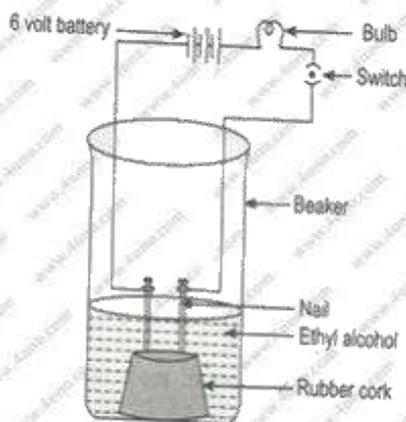
(b) How does the concentration of hydrogen ions $[H_3O]^+$ changes when the solution of an acid is diluted with water?

(c) Which has higher pH. A concentrated or dilute solution of HCL?

(d) What would you observe on adding dil HCL acid to

- (i) Sodium bicarbonate placed in a test tube.
- (ii) Zinc metal in a test tube.

Q.26. Compound such as alcohols and glucose also contain hydrogen but are not categorized as acids. Describe an activity



Q.27. State reason for the following statements:

- (i) Tap water conducts electricity whereas distilled water does not.**
- (ii) Dry hydrogen chloride gas does not turn blue litmus red whereas dilute hydrochloric acid does.**
- (iii) During summer season, a milk man usually adds a very small amount of baking soda to fresh milk.**
- (iv) For a dilution of acid, acid is added into water and not water into acid.**
- (v) Ammonia is a base but does not contain hydroxyl group.**

Q.28. . A substance 'X' used in the kitchen for making tasty crispy pakoras and is also an Ingredient of antacid. Name the substance 'X'.

- (i) How does 'X' help to make cakes and bread soft and spongy.**
- (ii) Is the pH value of solution of 'X' is lesser than or greater than 7.0?**

Q.29. Write word equations and then balanced equations for the reaction taking place when:

- (a) Dilute Sulphuric acid reacts with zinc granules.**
- (b) Dilute hydrochloric acid reacts with magnesium ribbon.**
- (c) Dilute Sulphuric acid reacts with aluminum powder**
- (d) Dilute hydrochloric acid reacts with iron fillings.**

Q.30. A road tanker carrying an acid was involved in an accident and its contents spilled on the road. At the side of the road iron drain cover began melting and fizzing as the acid ran over them. A specialist was called to see if the acid actually leaked into the nearby river.

- (a) Explain why specialist could carry out sample test to see of the river water contains some acid or not**
- (b) Suggest a better report name for the word 'melting'**
- (c) Explain why the drain covers began fizzing as the acid ran over them.**

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