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CBSE 10th Chemical reactions and Equations Unsolved Paper

SECTION - A

Q.1. What happens when dilute hydrochloric acid is added to iron filling? Tick the correct answer

- (a) Hydrogen gas and iron chloride are produced.
- (b) Chlorine gas and iron hydroxide are produced
- (c) No reaction takes place
- (d) Iron salt and water are produced

Q.2. Which of the following is not a balanced equation?

- (a) $Fe + Cl_2 \rightarrow FeCl_2$
- (b) $Mg + CuSO_4 \rightarrow MgSO_4 + C_4$
- (c) $NaOH + HCl \rightarrow NaCl + H_2O$
- (d) $Zn + S \rightarrow FeCl_3$

Q.3. Dissolving sugar is an example of-

- (a) Physical change
- (b) Chemical change
- (c) Redox Reaction
- (d) None of these.

Q.4. Take about 5 ml of dil. HCl in a test tube and add a few pieces of fine granules to it. Which gas is evolved?

- (a) Chlorine
- (b) Hydrogen
- (c) HCl
- (d) Nitrogen

Q.5. A substance which oxidizes itself and reduces other is known as

- (a) Oxidising agent
- (b) reducing agent
- (c) Both (a) and (b)
- (d) None of these.

Q.6. In the reaction $PbO + C \rightarrow Pb + CO$

- (a) Pbo is oxidised
- (b) C act as an oxidising agent
- (c) C act as a reduction agent
- (d) Reaction does not represent redox reaction.

Q.7. Copper displaces which of the following metals from its salt solution:

- (a) $ZnSO_4$
- (b) $FeSO_4$
- (c) $AgNO_3$
- (d) $NiSO_4$

Q.8. Which of the following reactions is not correct

- (a) $Zn + CuSO_4 \rightarrow ZnSO_4 + Cu$
- (b) $2Ag + Cu(NO_3)_2 \rightarrow 2AgNO_3 + Cu$
- (c) $Fe + CuSO_4 \rightarrow FeSO_4 + Cu$
- (d) $Mg + 2HCl \rightarrow MgCl_2 + H_2$

Q.9. PbS reacts with ozone (O_3) and forms $PbSO_4$. As per the balanced equation, molecules of ozone required for every one molecule of PbS is / are

- (a) 4
- (b) 3
- (c) 2
- (d) 1

Q.10. Some crystals of copper sulphate were dissolved in water. The colour of the solution obtained would be

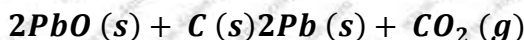
- (a) green
- (b) red
- (c) blue
- (d) brown.

Q.11. A shiny brown coloured element 'X' on heating in air becomes black in colour. Name the element 'X' and the black coloured compound formed.

Q.12. What do you mean by a precipitation reaction? Explain by giving examples.

Q.13. In refining of silver, the recovery of silver from silver nitrate solution involved displacement by copper metal. Write down the reaction involved.

Q.14. Which of the following statement about the reaction below are incorrect?



- (a) Lead is getting reduced.
- (b) Carbon dioxide is getting oxidized
- (c) Lead oxide is getting oxidized
- (d) Lead is getting reduced

- i. (a) and (b)
- ii. (a) and (c)
- iii. (a), (b) and (c)
- iv. All

Q.15. A solution of a substance 'X' is used for white washing

- i. Name the substance 'X' and writes its formula.
- ii. Write the reaction of the substance 'X' named in (i) above with water

Q.16. A student burnt a metal A found in the form of ribbon. The ribbon burnt with a dazzling Flame & a white powder B is formed which is basic in nature. Identify A & B Write the Balanced chemical equation.

Q.17. A substance X used for coating iron articles is added to a blue solution of a reddish brown metal Y, the color of the solution gets discharged. Identify X and Y & also the type of reaction.

Q.18. Why respiration is considered an exothermic reaction? Explain.

Q.19. What is a balanced chemical equation? Why should a chemical equation be balanced?

Q.21. Write the balanced equation for the following reactions. Give reasons for the following reactions?

i. Hydrogen + Chlorine → Hydrogen chloride

ii. Barium chloride + Aluminium sulphate

→ Barium sulphate + Aluminium chloride

iii. Sodium + water → Sodium hydroxide + water

Q.22. Astha has been collecting silver coins and copper coins. One day she observed a black coating on silver coins and a green coating on copper coins. Which chemical phenomenon is responsible for these coatings? Write the chemical name of black and green coatings?

Q.23. Name the type of reaction seen in the diagram below. Write the reaction for the same.

Q.24. Transfer the following into chemical equations and balance them.

(i) Hydrogen gas combines with nitrogen to form ammonia.

(ii) Hydrogen sulphide gas burns in air to give water and sulphur dioxide.

(iii) Potassium metal reacts with water to give potassium hydroxide and hydrogen gas.

Q.25. Explain the following terms with one example each.

a. Corrosion

b. Rancidity.

Q.26. Explain the following in terms of gain and loss of oxygen with two examples each?

a. Oxidation

b. Reduction

Q.27. What is difference between displacement and double displacement reactions? Write equations for these reactions.

Q.28. What does one mean by exothermic and endothermic reactions? Give examples.

Q.29. Write the balanced chemical equation for the following and identify the type of reaction in each case.

(a) Potassium bromide (s) + Barium iodide (aq)

→ Potassium iodide (aq) + Barium bromide(s)

(b) Zinc carbonate (s) → Zinc oxide (s) + Carbon dioxide (g)

(c) Hydrogen (g) + Chlorine (g) → Hydrogen chloride (g)

(d) Magnesium (s) + Hydrochloric acid (aq)

→ Magnesium chloride (aq) + Hydrogen (g)

Q.30. Write the balanced chemical equations for the following reactions.

(a) Calcium hydroxide + Carbon dioxide

→ Calcium carbonate + Water

(b) Zinc + Silver nitrate → Zinc nitrate + Silver

(c) Aluminum + Copper chloride → Aluminum chloride + Copper

(d) Barium chloride + Potassium sulphate

→ Barium sulphate + potassium chloride

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