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# CBSE 12th Chemistry 2008 Unsolved Paper Delhi Board

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# CBSE 12th Chemistry 2008 Unsolved Paper Delhi Board

TIME - 3HR. | QUESTIONS - 30

THE MARKS ARE MENTIONED ON EACH QUESTION

## SECTION – A

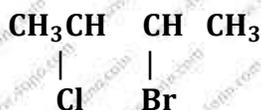
**Q. 1. What is the coordination number of each type of ions in a rock-salt type crystal structure? 1 mark**

**Q. 2. Define the term 'order of reaction' for chemical reactions. 1 mark**

**Q. 3. What causes Brownian movement in a colloidal solution? 1 mark**

**Q.4. In which one of the two structures,  $\text{NO}^+_2$  and  $\text{NO}^-_2$ , the bond angle has a higher value? 1 mark**

**Q.5. Write the IUPAC name of the following compound: 1 mark**



**Q.6. Arrange the following compounds in an increasing order of their acid strengths:  $(\text{CH}_3)_2\text{CHCOOH}$ ,  $\text{CH}_3\text{CH}_2\text{CH}(\text{Br})\text{COOH}$ ,  $\text{CH}_3\text{CH}(\text{Br})\text{CH}_2\text{COOH}$ . 1 mark**

**Q.7. Write a chemical reaction in which the iodide ion replaces the diazonium group in a diazonium salt. 1 mark**

**Q.8. Name a substance that can be used as an antiseptic as well as a disinfectant. 1 mark**

## SECTION – B

**Q.9. Explain as to why haloarenes are much less reactive than halo-alkanes towards nucleophilic substitution reactions. 2 marks**



- Q. 20.** Calculate the temperature at which a solution containing 54g of glucose, ( $C_6H_{12}O_6$ ), in 250g of water will freeze. ( $K_f$  for water =  $1.86 \text{ K mol}^{-1} \text{ kg}$ ) 3 marks
- Q. 21.** What are lyophilic and lyophobic sols? Give one example of each type. Which one of these two types of sols is easily coagulated and why? 3 marks
- Q. 22.** State briefly the principles which serve as basis for the following operations in metallurgy: 3 marks
- Froth floatation process
  - Zone refining
  - Refining by liquation
- Q. 23.** Write chemical equations for the following processes: 3 marks
- Chlorine reacts with a hot concentrated solution of sodium hydroxide
  - Orthophosphorous acid is heated
  - Pt  $F_6$  and xenon are mixed together

Or

Complete the following chemical equations:

- $Ca_3P_2(s) + H_2O(l) \rightarrow \dots$
  - $Cu^{2+}(aq) + NH_3(aq) \rightarrow \dots$   
(excess)
  - $F_2(g) + H_2O(l) \rightarrow \dots$
- Q.24.** (a) What is a ligand? Give an example of a bidentate ligand. 3 marks  
(b) Explain as to how the two complexes of nickel,  $[Ni(CN)_4]^{2-}$  and  $Ni(CO)_4$  have different structures but do not differ in their magnetic behavior. (Ni = 28)
- Q. 25.** Name the reagents which are used in the following conversions: 3 marks
- A primary alcohol to an aldehyde
  - Butan-2-one to butan-2-ol
  - Phenol to 2, 4, 6-tribromophenol
- Q. 26.** Account for the following observations: 3 marks
- $pK_b$  for aniline is more than that for methylamine.
  - Methylamine solution in water reacts with ferric chloride solution to give a precipitate of ferric hydroxide.
  - Aniline does not undergo Friedel-Crafts reaction.

- Q. 27.** Write the names and structures of the monomers of the following polymers 3 marks
- Buna-S
  - Neoprene
  - Nylon-6

SECTION – D

- Q. 28.** Conductivity of 0.00241 M acetic acid solution is  $7.896 \times 10^{-5} \text{ S cm}^{-1}$ . Calculate its molar Conductivity in this solution. If  $\Lambda_M^0$  for acetic acid be  $390.5 \text{ S cm}^2 \text{ mol}^{-1}$ , what would be its dissociation constant? 5 marks

